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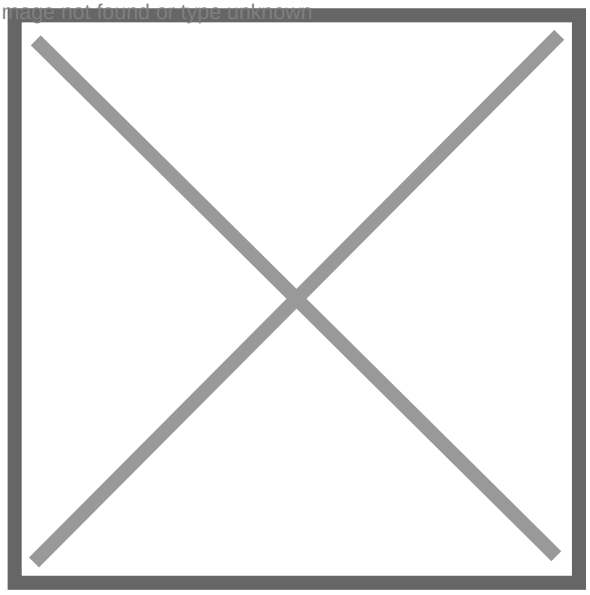
Product Name :

Molecular Model Set – Key Stage 3 Chemistry

Product Code :

SLE/Mo/31

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Description :

Molecular Model Set – Key Stage 3 Chemistry (Product Code: SLE/Mo/31) is a molecular model kit designed for the Key Stage 3 Chemistry syllabus, used to model atoms, elements, compounds and chemical reactions in lower secondary science. It uses colour-coded parts and places its emphasis on showing conservation of mass, with reactions that range from neutralisation and rusting to combustion and thermal decomposition. It is made by **Jain Scientific Equipments Pvt Ltd (ScienceLab)**, an OEM manufacturer, at its own factory in Ambala, India.

Key Details

Parameter	Detail
Syllabus	Designed for use with the Key Stage 3 Chemistry syllabus
Concepts introduced	Atoms, elements, compounds, chemical symbols and molecular formulae
Representation	Colour-coded parts
Teaching emphasis	Modelling the conservation of mass during chemical reactions
Product code	SLE/Mo/31
Manufacturer	Jain Scientific Equipments Pvt Ltd (ScienceLab) — OEM
Country of origin	India (made in Ambala, Haryana)

The set can be used to model these reactions and examples:

- Displacement reactions of metals in salts
- Examples of common acids, bases and alkalis
- Extraction of metals by the reduction of metal oxides with carbon
- Neutralisation reactions of acids and alkalis
- Material examples of polymers, PVC and polyethylene
- The greenhouse effect through the production of CO2
- Reactions of acids with metals to produce a salt and hydrogen
- Reactions of acids and alkalis to produce a salt and water
- Reactions of acid rain on soil, metals and limestone
- The chemical reaction of respiration on glucose
- Oxidation by combustion of hydrocarbons
- Oxidation through the rusting of metals
- Thermal decomposition of metal carbonates

What Is the Key Stage 3 Chemistry Model Set Used For?

Lower secondary students often find word and symbol equations abstract. With this kit they build the reactant molecules, rearrange the same parts into products and count the atoms on each side, which shows that nothing is lost or created in a reaction. The same approach works for acid and alkali neutralisation, metal extraction with carbon and the combustion of hydrocarbons.

Link the models to everyday contexts in the syllabus: carbon dioxide and the greenhouse effect, acid r
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