Email : sales@scilabexport.com
Phone: +91-7082934803

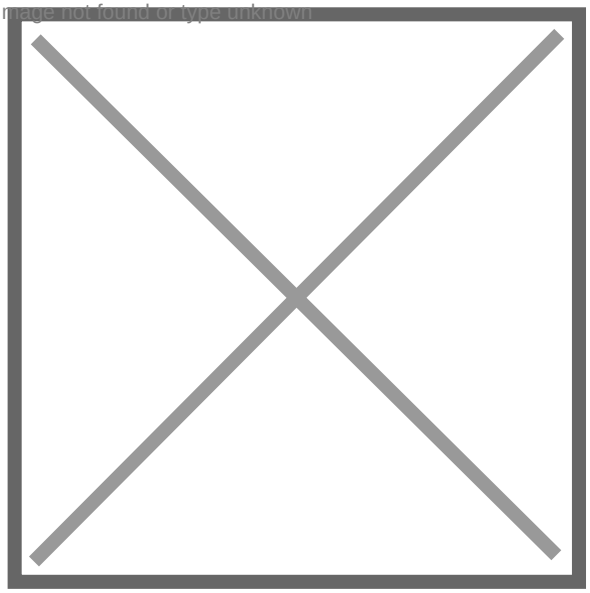
Product Name :

Hydrochloric Acid Cone. C.P grade

Product Code :

SCL-867/176

Image not found or type unknown



Email : sales@scilabexport.com
Phone: +91-7082934803

Description :

Hydrochloric Acid Cone. C.P grade (Product Code: SCL-867/176) is concentrated hydrochloric acid used as the stock acid from which laboratory staff prepare dilute bench solutions in school and college chemistry labs. The 500 ml bottle is meant for teachers and technicians who keep concentrated acid locked away; its name gives chemically pure (C.P) grade, while the listing also calls it an analytical reagent. It is sent with an MSDS and supplied by **Jain Scientific Equipments Pvt Ltd (ScienceLab)**, Ambala, India.

Specifications

Parameter	Detail
Chemical	Hydrochloric acid (HCl), concentrated
Grade	C.P grade as named; the listing also describes it as an analytical reagent – confirm the grade you need when ordering
Capacity	500 ml
Safety documentation	Supplied with MSDS
Assay	Not stated in this listing – ask us if you need it
Shelf life	Shelf life depends on the chemical, its grade and storageconditions; ask us for the expiry date of the batch you are quoted.
Product code	SCL-867/176
Supplied by	Jain Scientific Equipments Pvt Ltd (ScienceLab), Ambala, India

What Is Concentrated Hydrochloric Acid Used For?

Most of the bottle goes into making dilute hydrochloric acid for the bench: staff measure the water first and add the concentrated acid slowly with stirring, in the ratio set by the practical manual. Used directly, a glass rod dipped in the acid and held near ammonia gas produces dense white fumes of ammonium chloride, the standard confirmatory test for ammonia and ammonium salts.

For flame tests, a nichrome or platinum wire is cleaned by dipping it in the acid and heating until the flame shows no colour, and the test salt is moistened with a drop of acid to give a stronger flame colour. The acid releases corrosive hydrogen chloride fumes as soon as the cap is removed, so open and dilute it only inside a fume cupboard, wearing goggles, acid-resistant gloves and an apron.

Store the bottle in a lockable chemistry reagent cabinet, away from alkalis, metal shelving and oxidisers such as bleaching powder or potassium permanganate, which release chlorine with it. Keep an acid spill kit within reach and never leave the bottle where students can reach it unsupervised.

Supplier

Jain
Scien
and c
intern